

Endothermic and Exothermic Reactions

~~In~~ During a chemical reactions, reactants are converted into products by the making and breaking of chemical bonds ~~converts the reactants into products.~~ When two or much more substances ~~counter~~ react, chemical bonds ~~are~~ is formed among between atoms, ~~which~~ creating create and forms a new chemical compounds. All chemical reactions are accompanied by a change in energy and can be ~~Chemical reactions are~~ classified into as either two types: endothermic and or exothermic reactions. In addition, activation ~~E~~ energy activation results in the ~~is~~ required to bonding of the two reactants to form a new product. ~~All chemical reactions are accompanied by a change in energy.~~

Comment [A1]: "During" indicates a period or range of time (having duration) and is used to say that something happened. A clause with during focuses more on what happened—the activity, event, or experience. A clause with "in" focuses more on when something happened rather than what happened.

~~There are many chemical~~ Exothermic reactions involve the emit release of energy with forms in the form of heat, light, or sound. ~~Such chemical reactions are called exothermal reactions.~~ This release of energy ~~that is released comes from~~ caused by the bonds that join bonding of several atoms ~~together~~ in the molecules participating in the reaction. ~~A~~ Combustion is a common example of ~~exothermal~~ exothermic reactions ~~is the phenomena of.~~ Complete combustion. ~~A complete combustion process is a~~ occurs when a compound reacts with an oxidizing agent ~~substance, and the,~~ yielding compounds of each element in the fuel with the oxidizing element agent ~~a re~~ emitted. ~~There are mostly as products.~~ Most exothermic reactions are spontaneous ~~exothermal processes.~~ ~~On the other hand, many chemical,~~

Comment [A2]: Two sentences have been combined and simplified to bring out the essential point at this instance.

Conversely, endothermic reactions ~~absorb~~ involve absorption of energy in the form of heat, light, or sound ~~forms.~~ ~~Such chemical reactions are called endothermal reactions.~~ These reactions cannot ~~progress with no~~ begin or proceed without the addition of ~~heat or supplying~~ energy. ~~The resulting~~ ~~product of the~~ reaction ~~has lesser stability~~ products are less stable because, ~~the higher~~ the stability of a molecule decreases as the energy ~~bond, the less strength~~ of its ~~molecules possess.~~ A constituent bonds

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increases. Photosynthesis is a common example of ~~endothermal~~ endothermic reactions ~~is, which involves~~ the ~~phenomena of photosynthesis~~. Here, ~~plants~~ use ~~the~~ of energy from the sun to convert carbon dioxide and water into glucose and oxygen. Most ~~endothermal~~ endothermic reactions are ~~not non-~~spontaneous.

Comment [A3]: The sentence is revised for clarity. Note that the statement seems contrary to the general knowledge.

To understand the difference between the two ~~reactions~~ types of reactions, we need to explore several concepts ~~like, such as~~ the behavior of kinetic ~~energy~~ and potential ~~energy behavior in the molecules of~~ the reactants of the chemical reaction energies of the reactant molecules.

Comment [A4]: "Like" is used when comparing things that have similar qualities, quantities, or degree. "Such as" is used when talking about specific things.

Comment [A5]: The original sentence was unclear and needed complete rewriting to make the sentence unambiguous. Redundancies ("of the reactants of the chemical reaction") have been removed and appropriate words ("to understand" instead of "to know") have been used.

SAMPLE